The Activities of Strong Electrolytes. III. The Use of the Flowing Junction to Study the Liquid-junction Potential between Dilute Hydrochloric Acid and Saturated Potassium Chloride Solutions; and the Revision of some Single-electrode Potentials, by George Scatchard.

P. 704, line 30, p. 705, line 1 and p. 708, line 28, instead of "3.5," read "0.35."

P. 705. Fig. 4, instead of the ordinates "1, 2, 3, 4 mv.," read "0.1, 0.2, 0.3, 0.4 mv."

The Concentration of Ammonia in a Compressed Mixture of Hydrogen and Nitrogen over Liquid Ammonia, by Alfred T. Larson and Charles A. Black.

P. 1018. In line 5 of Section 3, instead of "The compressibility of liquid ammonia has been determined by Cragoe and Harper,[6]" read "The specific volume of liquid ammonia has been determined by Cragoe and Harper,[6] and its compressibility is known."[6a]

Insert at the bottom of the page Ref. 6 a, thus:

[6a] "Tables of Thermodynamic Properties of Ammonia," *Bur. Standards Circ.*, **142**, 30 (1923).

The Quantitative Determination of Vitamin A, by H. C. Sherman and H. E. Munsell.

P. 1641. In Ref. 8, instead of "THIS JOURNAL," read "*J. Biol. Chem.*"

Displacement of Metals from Solutions of Their Salts by Less Electropositive Elements. II. The Reaction between Amides of the Alkali and Alkaline Earth Metals and Elements More Electropositive than Tin, by F. W. Bergstrom.

P. 1837. Line 18, instead of "Table II," read "Table IV."

P. 1838. Line 9, instead of "other ammono salts," read "potassium salts of the metallic imides and amides—."

P. 1839. Line 16, instead of "Equation 3," read "Equation 1."

The Catalytic Decomposition of Hydrogen Peroxide in an Acid Chlorine-Chloride Solution, by Robert S. Livingston and William C. Bray.

P. 2080. A note correcting some of the equations and calculations will appear in a later number of THIS JOURNAL.

Alkyl-allyl-barbituric Acids, by Ernest H. Volwiler.

P. 2239. In Table III, heading of Col. 2, instead of "inner ear," read "outer ear passage."

Studies on the Directive Influence of Substituents in the Benzene Ring. III. The Active Agent in Aqueous Bromination, by A. W. Francis.

P. 2347. In the third line from the bottom, omit "with."

The Absorption Spectra of Benzaurin, by W. R. Orndorff, R. C. Gibbs and Alice McNulty.

Pp. 2774 and 2775. The printed legends and captions referring to Fig. 4, including the figure number, belong with the cut printed as Fig. 5, and those printed with Fig. 5 belong with the cut printed as Fig. 4.

NEW BOOKS

The Structural Units of the Material Universe. Seventh Earl Grey Memorial Lecture, delivered at King's Hall, Armstrong College, Newcastle-on-Tyne, March 5, 1925. By F. W. ASTON, Sc.D. Oxford University Press, American Branch, New York, 1925. 23 pp. 5 figs. 15 × 23 cm. Price $0.35.

This is a simple and straightforward account, by one of the protagonists, of the development of our knowledge of the structure of atoms, culminating in the mass-spectrograph and the artificial disintegration of atoms.

ARTHUR B. LAMB

The Pronouncing Chemical Formula Speller and Contest Guide. By C. A. JACOBSON, Professor of Chemistry, West Virginia University. The Williams and Wilkins Company, Baltimore, U. S. A., 1925. xviii + 279 pp. 20.5 × 14 cm. Price $4.00.

A teacher of organic chemistry used to speak of nitric acid (long i) and nitrobenzene (short i). When questioned, he stated each was correct because he and others pronounced them so. Why a colleague of his pronounced one of them differently, he did not explain.

Any influence tending to bring uniformity in chemical pronunciation is to be welcomed. When this carries with it the joy and the zest of a contest, it is doubly effective.

Professor Jacobson has been a leader in the recent development of the chemical spelling match and his book is a real contribution. His painstaking care and the attention given to detail are evident on every page.

The list of inorganic compounds occupies 78 pages, that of minerals about 10 pages, and of the organic compounds about 46 pages. More important compounds are in capitals. In all of these lists pronunciation is shown by diacritical marks. Part IV, of typical chemical reactions for each element arranged alphabetically, occupies the latter half of the book.

Even in so comprehensive a list there are some omissions. Carbamic acid is listed, but not carbonic; likewise, borofluoric, but not boric. For many of the organic compounds two names are given; thus, sucrose and cane sugar.

The book is clearly intended not only as a guide for chemical contests, but as an effort to standardize chemical nomenclature. For the first purpose it is ideal except, perhaps, for its comprehensiveness. The number of formulas given might tend to discourage the weaker student and forestall effort, though stimulating the more capable. This very fact makes the book invaluable for reference in every chemical laboratory, whether a contest is to be held or not. That such a contest can equal a base-ball game in interest for both contestants and spectators, we know by experience.

For the standardization of chemical nomenclature and the securing of uniform pronunciation the book gives much promise. The American Chemical Society might well refer this matter to a regular or a special committee, who could use this list as a starting point. Perhaps such a committee could succeed in eliminating some of the nine scunds given to the single letter, a, and also to o, according to present diacritical markings. Surely they could improve a situation where such frequently used words as chlorine, phenolphthalein, amide, have two pronunciations. We might finally all agree to make pyrites singular, having it pyrite, ending in -ite as most good minerals do.

At all events we are indebted to Professor Jacobson for a useful book of decided value.

HARRISON HALE